Is the following reaction a reduction, an oxidation, or neither an oxidation or reduction reaction. Show your reasoning.

Given the following:

A(triphosphate) + glucose \Leftrightarrow A(diphosphate) + glucose-6-phosphate ($\Delta G^o = -4.0 \text{ kcal mol}^{-1}$) A(triphosphate) \Leftrightarrow A(diphosphate) + phosphate ($\Delta G^o = -7.3 \text{ kcal mol}^{-1}$)

What is the ΔG^o for: glucose + phosphate \Leftrightarrow glucose-6-phosphate

At equilibrium are more products or reactants present?

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